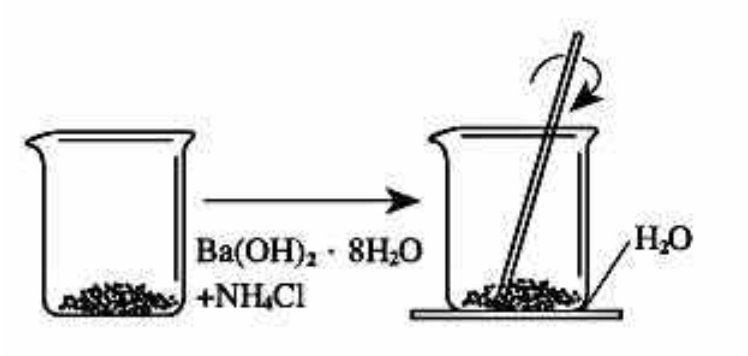
Endothermic vs Exothermic

Objective: I will be able to determine when energy is lost or gained in a system.

Two solids get mixed together in a beaker:

$$Ba(OH)_2 + NH_4Cl \rightarrow BaCl_2 + NH_3$$

Set the beaker on a few drops of water.



1. How does the reaction feel?	Cold.
2. Why does the H_2O under the beaker freeze?	The chemical reaction absorbed heat from the water, causing the water to freeze.
Sublimation = solid to gas ⇄ Deposition = gas to solid Solid ← → Liquid ← → Gas Low/less energy ← → High/more energy	
3. System	The main object
4. Surroundings	Everything else around it.